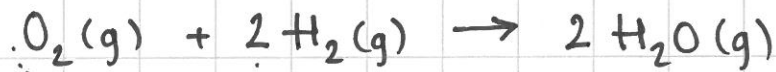
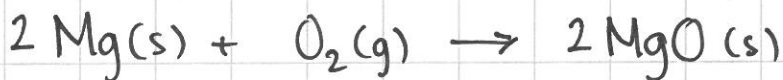


TASAPAINOTETUN REAKTION YHTÄLÖN KÄYTTÖ



$$\frac{n(\text{O}_2)}{n(\text{H}_2)} = \frac{1}{2}$$

ESIM $m(\text{Mg}) = 100\text{mg}$, palaa
 $m(\text{MgO}) = ?$



$$n(\text{Mg}) = \frac{m}{M} = \frac{0,100\text{g}}{24,31\text{g/mol}} \approx 4,1135 \cdot 10^{-3}\text{mol}$$

$$\text{sama kerroin} \Rightarrow n(\text{Mg}) = n(\text{MgO})$$

$$n(\text{MgO}) = 4,1135 \cdot 10^{-3}\text{mol}$$

$$M(\text{MgO}) = (24,31 + 16,00)\text{g/mol} = 40,31\text{g/mol}$$

$$m(\text{MgO}) = n \cdot M = 0,16582\text{g} \approx 166\text{mg}$$

$$n = \frac{m}{M}$$

reaktioyhtälön
kerroin

$$\frac{n(\text{A})}{n(\text{B})} = \frac{\dots}{\dots}$$

$$m = n \cdot M$$

1.7

1.8

1.9

1.11

1.14