

$$\begin{aligned} \underline{3.31} \\ \text{eskin} \rightarrow \begin{cases} m(\text{alkoholi}) = 0,720 \text{ mg} \\ m(\text{etiilihappo}) = 0,400 \text{ mg} \end{cases} \end{aligned}$$

$$M_r(\text{alkoholi}) \approx 108$$

$$\text{ps } m(\text{alkoholi}) = 1,00 \text{ g}$$

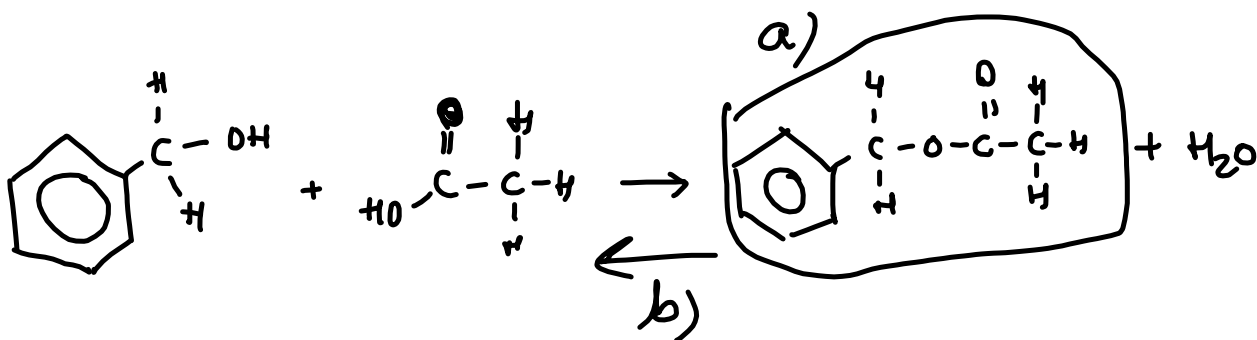
$$\begin{cases} m\%(\text{C}) = 77,78\% \\ m\%(\text{H}) = 7,41\% \\ m\%(\text{O}) = 14,81\% \end{cases}$$

$$\begin{cases} n(\text{C}) = 6,476 \text{ mol} \\ n(\text{H}) = 7,351 \text{ mol} \\ n(\text{O}) = 0,926 \text{ mol} \end{cases} \quad (0,926)$$

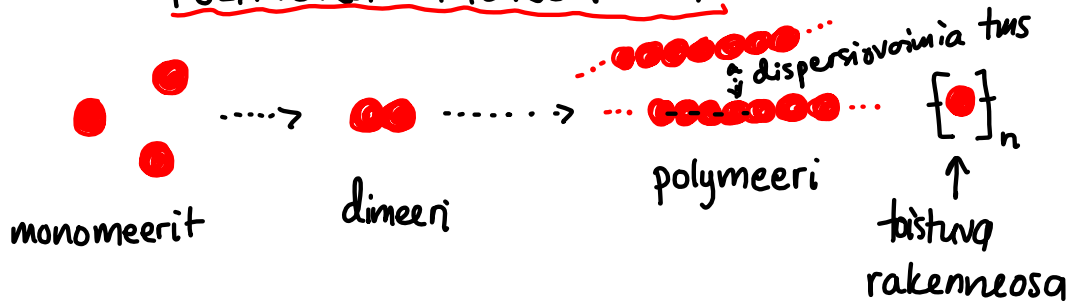
$$n(\text{C}) : n(\text{H}) : n(\text{O})$$

$$\approx 7,0 : 7,9 : 1 \approx 7:8:1$$

$$M(\text{C}_7\text{H}_8\text{O}) = 108,134 \text{ joten molekyylikaava}$$



4.1 POLYMEERIT JA POLYMEROITUMISREAKTIOT

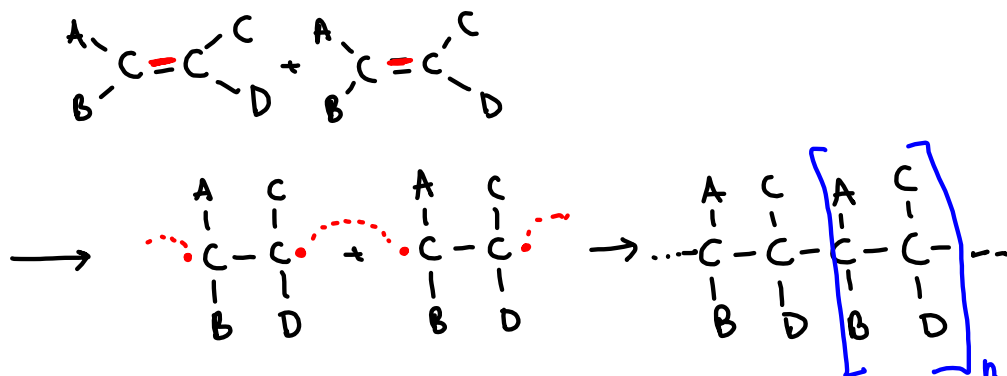


polykondensaatio:

monomeerissa kaksi kondensaatioreaktion osallistuvaa funktionaalista ryhmää ($-OH$, $-COOH$, $-NH_2$)
→ polyesterit, polyamidit (Nylon)

polyadditio:

monomeerissa kaksoissidos, joka aukeaa



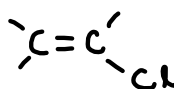
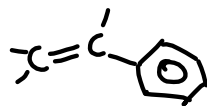
polyeteeni

polypropeeni

polystyreeni

polyvinyylikloridi

teflon



PVC

4.2

4.3

4.7

4.7

Nailon 66

